

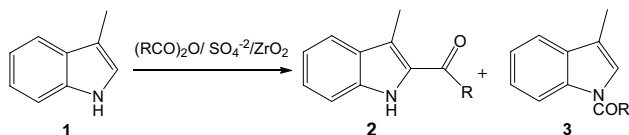
Sulfated Zirconia: An Efficient and Reusable Heterogeneous Catalyst in the Friedel–Crafts Acylation Reaction of 3-Methylindole

Darío A. Vargas¹ · Leticia J. Méndez^{1,2} · Alicia S. Cánepa¹ · Rodolfo D. Bravo¹

Received: 20 January 2017 / Accepted: 17 April 2017 / Published online: 27 April 2017
© Springer Science+Business Media New York 2017

Abstract An efficient process for the Friedel–Crafts acylation of 3-methylindole with acid anhydrides in the presence of sulfated zirconia was developed. This catalyst shows excellent catalytic activity with good conversion and selectivity towards formation of the products of 2-acylation (3-methyl-1*H*-indol-2-yl)ketones. Other advantages related to this process are the simple work-up procedure and smaller production of chemical waste. The catalyst was easily recycled and reused with a minimal loss in activity through three reaction cycles. The catalyst was characterized by FT-IR spectroscopy, X-Ray diffraction and superficial acidity.

Graphical Abstract



Keywords Sulfated zirconia · 3-Methylindole · Ketones · Recyclability · Acylation reaction

1 Introduction

Indole derivatives present wide biological importance in a variety of therapeutic areas such as anti-inflammatory, anticonvulsant, cardiovascular, antitumor, antibacterial, antiparkinsonian, antiviral and antidiabetic agents, among others [1].

In particular, acylindoles present appreciable interest, not only for their biological properties but they are also widely used as intermediates in the synthesis of alkaloid and different families of heterocyclic compounds [2, 3].

There are different methods in the synthesis of acylindoles, among them the Friedel–Crafts reactions [4–7], Vilsmeier–Haack reaction [8] and the use of Grignard reactions [9].

Some of these methods present advantages and shortcomings that limit their scope and performance. The Vilsmeier–Haack acylations also give good yields, but the amides used are limited (e.g. formamide, alkylcarboxamide) and a large amount of POCl₃ is used, causing complications in the environment.

Regarding the Friedel–Crafts reaction, the electrophilic substitution preferably occurs at the C-3 instead of the C-2 position. Acylation at the C-2 position can occur for indoles 1,3-disubstituted or 3-substituted. However, in 3-substituted indole compounds, *N*-acylation may also occur [10]. The use of *N*-protecting groups is generally the chosen strategy to avoid the formation of 3-acyl derivatives [11]. The disadvantage of this process is the largest number of synthetic steps, with the initial preparation of the *N*-protected derivative, subsequent acylation and removal of the protecting group to give the final product.

Our research group is interested in obtaining 2-acyl-3-methylindoles, which are used as substrates in the synthesis of complex heterocyclic compounds. Thus,

✉ Alicia S. Cánepa
ascanepa@quimica.unlp.edu.ar

¹ Centro de Estudio de Compuestos Orgánicos, Departamento de Química, Facultad de Ciencias Exactas, Universidad Nacional de La Plata, Calle 47 y 115, 1900 La Plata, Argentina

² CINDECA (CONICET-UNLP), Facultad de Ciencias Exactas, 47 No. 257, 1900 La Plata, Argentina

Friedel–Crafts acylation arises as a good synthetic way of mentioned compounds from 3-methylindole, using Lewis or Brønsted acid catalysts. Homogeneous Lewis acids used for these reactions include AlCl_3 , BF_3 , TiCl_4 , ZnCl_2 , SnCl_2 , SnCl_4 and ZrCl_4 among others [7, 12–14].

These days, a number of factors demand the use of heterogeneous catalysts because of problems associated with the handling, disposal of homogeneous acids and environmental contamination.

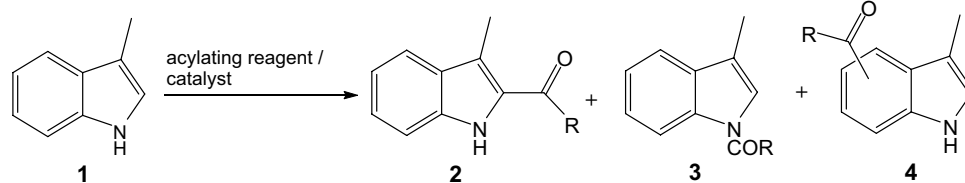
We have studied the synthesis of different compounds through the use of homogeneous or heterogeneous acid catalysts. So, sulfuric acid, methanesulfonic acid (MSA) and trifluoroacetic acid (TFA) were used as homogeneous catalysts [15, 16]. Moreover, many of these processes were performed using heterogeneous catalysts to compare their viability. The catalysts used were tungsten and molybdenum heteropolyacids ($\text{H}_3\text{PWO}_{40}$ and $\text{H}_3\text{PMoO}_{40}$) supported on silica [17], Amberlyst 15 and Amberlyst XN-1010 resins [15, 18–21], sulfated zirconia [22, 23], P_2O_5 and HClO_4 supported on silica [24]. In almost all cases, yields increased when heterogeneous catalysts were used, with simplicity in the purification process and a smaller amount of polluting waste.

Among those mentioned sulfated zirconia has attracted much attention in recent years because of its good catalytic activity, super-acidity, non-toxicity and several advantages such as short reaction times, high selectivity and the easiness of work-up procedure. Some interesting methods using this catalyst include Friedel–Crafts acylation [25], esterification reactions [26–28], synthesis of flavones [29], synthesis of coumarins by the Pechmann reaction [30] and 1,5-benzodiazepines [31].

Continuing with our work on the application of heterogeneous catalysts for development of useful synthetic methods, we decided to study the acylation of 3-methylindole using sulfated zirconia as a catalyst.

Referring to the Friedel–Crafts acylation of indole derivatives, mono- and bi-substituted products can be obtained. In relation to obtaining mono acylated products, these may be substituted at the C-2 position (**2**), at the *N*-position (**3**) and other carbon atoms in the benzene ring (**4**) (Scheme 1).

Scheme 1 Possible products in mono acylation of 3-methylindole



2 Experimental

2.1 Materials

Melting points were determined with a Buchi apparatus. ^1H NMR and ^{13}C NMR spectra were recorded on a Bruker AVIII 600 Biospin in CDCl_3 or DMSO-d_6 . Thin layer chromatography was performed on silica gel sheets 60 F₂₅₄ (Merck A.G.). Silica gel 60 (70–230 mesh) (Fluka) was used for column chromatography. Commercial 3-methylindole (98%) (Aldrich) was recrystallized from hexane before use. All reagents were purchased from Merck A.G. and Aldrich; 1,2-dichloroethane and 1,1,2-trichloroethane were distilled over phosphorous pentoxide and stored over 4 Å molecular sieves. Commercial $\text{SO}_4^{2-}/\text{ZrO}_2$ from Mel Chemical Co was dried at 110 °C for 2 h in an air and was subsequently calcined for 5 h in air atmosphere at 550 °C, based on previous experiences [21]. All yields refer to isolated products.

2.2 Catalyst Characterization

The solids were characterized by FT-IR spectroscopy, X-Ray diffraction (XRD) and superficial acidity.

X-ray diffractograms were obtained in a Philips PW-1732 equipment, using Cu K α radiation Ni filter; 20 mA and 40 kV as the high voltage source, scanning angle between 5° and 50° 2 θ , and scanning rate 1°/min. FT-IR spectra were obtained using a Bruker IFS 66 FTIR spectrometer.

The determination of total acid sites of the catalysts was measured by potentiometric titration with *n*-butylamine in acetonitrile as a non-aqueous medium. The solid (0.05 g) was suspended in acetonitrile, and stirred for 3 h. The suspension was titrated with a 0.05 N solution of *n*-butylamine in acetonitrile using a Metrohm 794 Titrino apparatus with a double junction electrode.

2.3 Reaction Study

Initially, the reaction between 3-methylindole (**1**) (1.0 mmol) and acetic anhydride as the acylating reagent (variable amounts) was investigated as a model to optimize the experimental conditions. Sulfated zirconia was used as a heterogeneous catalyst. Parallel reactions are performed

using acetic acid as an example of a homogeneous catalyst in order to compare their behavior (Scheme 1). These were carried out in a batch reactor with magnetic stirring under Ar atmosphere. The progress of the reaction was monitored by TLC (9:1 hexane/ethyl acetate). When the reaction was complete, the catalyst was separated by filtration, and the solvent was removed in vacuum. The crude products were purified by column chromatography on silica gel (9:1 hexane/ethyl acetate). The known reactions products were characterized by comparison of their physical and ^1H NMR and ^{13}C NMR spectra and data with those of reported ones. The identification of new compound was made through ^1H NMR and ^{13}C NMR spectra and elemental analysis.

Different reaction conditions were checked, catalyst amount, substrate/acylating reagent ratio and reaction time, for both type of catalysts.

The reactions were carried out either at 82°C using 1,2-dichloroethane as a solvent or at 140°C using 1,1,2,2-tetrachloroethane as a solvent.

2.4 Acylating Reagent

With the optimized conditions in hand, we probed the scope of acylation of 3-methylindole using other examples of acylating agents such as propionic anhydride, *n*-butyric anhydride and benzoic anhydride.

2.5 Reusability of Catalyst

The used catalysts were washed, according to the case, with 1,2-dichloroethane or 1,1,2,2-tetrachloroethane, then methanol and drying in air at 120°C for 2 h and reused in subsequent following experiments.

3 Results and Discussion

3.1 Catalyst Characterization

The properties of sulfated zirconia were examined by powder XRD, FT-IR spectroscopy and superficial acidity in catalysts calcined at 550°C , without use and after being used.

The XRD patterns of samples calcined at 550°C without use (a) and after being used (b) are shown in Fig. 1. As can be seen from Fig. 1 sulfated zirconia calcined at 550°C contained only tetragonal phases ($2\theta=30.16^\circ$, 34.96° , 50.22° and 59.98°). The catalyst recovered and reused shown the same pattern of bands, without changes after three catalytic cycles.

The total number and strength of Brønsted and Lewis acids sites were measured by nonaqueous potentiometric titration with *n*-butylamine in acetonitrile. In this method,

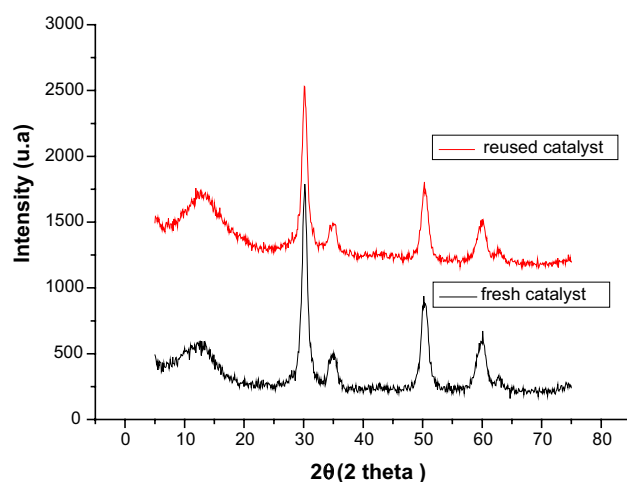


Fig. 1 XRD patterns of sulfated zirconia for fresh sample calcined at 550°C and a reused sample

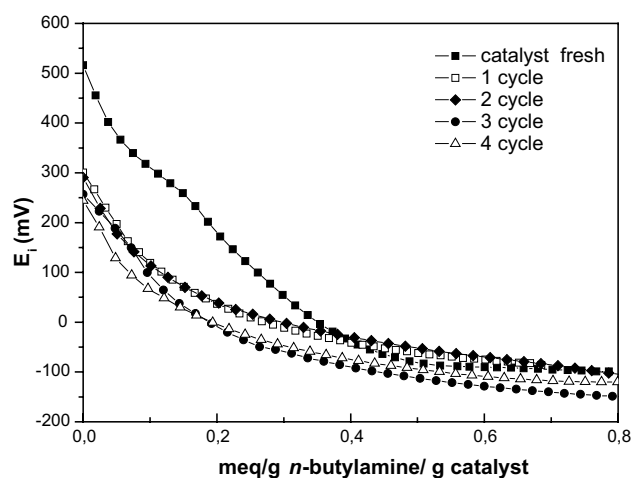


Fig. 2 Potentiometric titration of *n*-butylamine in acetonitrile for sample calcined at 550°C and after being used in different cycles

the initial electrode potential (E_i) indicates the maximum acid strength of the surface site, and the materials with E_i values higher than 100 mV are defined as very strong solid acids. Additionally, the range where the plateau was reached [$\text{meq}(n\text{-butylamine}) \text{g}^{-1}$] shows the [32, 33].

The measurements were performed on an original sample and after being used in different cycles for the acylation reaction using acetic anhydride as acylating reagent. The potentiometric titration curves are presented in Fig. 2.

It can be observed in the Fig. 2 that the initial electrode potential E_i is 516.2, 303.3, 290.6, 257 and 244.2 mV for the original sample of SO_4/ZrO_2 and samples used in four catalytic cycles respectively. These results indicate the presence of very strong acid sites.

The amount of *n*-butylamine consumed was 0.76 meq/g for the original sample of SO₄/ZrO₂ to 0.63, 0.80, 0.74 and 0.75 meq/g for SO₄/ZrO₂ after four cycles respectively.

The FT-IR spectra of catalysts calcined at 550 °C, without use and after being used are shown in Fig. 3. The FT-IR spectra shown bands of the SO₄²⁻ group in the region of 1200–900 cm⁻¹, with peaks at 999, 1041, 1072, 1142 and 1235 cm⁻¹, which are in agreement with those reported in the literature for sulfated zirconia assigned to asymmetric and symmetric stretching frequencies of ionized S=O double bonds and S–O bonds. An additional broad peak at 3408 cm⁻¹ corresponds to the stretching vibrations hydroxo- and aquo-OH of hydroxyl groups and adsorbed water accompanied by the band at 1628 cm⁻¹. In addition, the weak unresolved band between 800 and 520 cm⁻¹ is attributed to Zr–O stretching modes.

FT-IR spectra of reused zirconia sample washed and dried at 120 °C showed features similar to those observed for fresh sulfated zirconia. In addition, very low intensity bands at 1456, 2850 and 2927 cm⁻¹ can be assigned to the bending/scissoring and stretching vibration of C–H and CH₃ groups. The latter were generated, due to traces of adsorbed reagent.

3.2 Reaction Studies

3.2.1 Comparison of Homogeneous and Heterogeneous Acids as Catalyst in the Acetylation of 3-Methylindole

The key stage of this reaction mechanism involves the generation of acylium ion intermediate, obtained from the acylating reagent in the presence of a catalyst. As can be found described in literature, acylation reactions can be catalyzed by Brønsted or Lewis acid. On this occasion, acetic acid and sulfated zirconia were used as examples of

homogeneous and heterogeneous catalysts. As shown in Table 1 the best results were obtained by the use of sulfated zirconia. Two products were formed, generated mainly by the substitution at C-2 position (**2a**) and a lower yield of *N*-acylation product (**3a**) (Scheme 1).

3.2.2 Effect of Different Amounts of Sulfated Zirconia in the Acetylation of 3-Methylindole

The effect of different amounts of sulfated zirconia was studied in the acetylation of 3-methylindole in the range 25–100% (w/w). An increase in the yield of the reaction could be observed with the increase of the proportion of catalyst employed, obtaining the best results with 100% (w/w) of sulfated zirconia (Table 1).

3.2.3 Influence of the Temperature

To study the effect of temperature, the reaction was carried out at 82 and 140 °C. The results indicate better results at 140 °C using 1,1,2,2-tetrachloroethane as a solvent. At this temperature the products are formed in a short reaction time (0.25 h), greater yields and regioselectivity towards formation of product **2a**, substituted at the C-2 position (Table 1).

3.2.4 Influence of Molar Ratio of 3-Methylindole and Acylating reagent

The effect of the molar ratio between 3-methylindole and acylating reagent was studied in the range 1:2–1:4. Experiments showed that a ratio of 1:4 can be considered optimal to perform this process. As could be observed in Table 1, by the use a lower ratio of acylating reagent (1:2) the yields are lower with recovery of reagent in some cases.

3.2.5 Acylating Reagent

The reaction of 3-methylindole (**1**) using other anhydrides as acylating reagents were performed in order to study the selectivity in relation to the substituent. In this regard, reactions using anhydrides with alkyl substituent (R=Me, Et and *n*-propyl) showed limited selectivity with higher yields of C-2 acylated products, relative to the *N*-acylated products as shown in Table 2. In contrast, aromatic anhydride (R=Ph) provide a higher selectivity, with very good yield of C-2 acylated product (80%), without formation of another isomer. The stabilizing effect of the phenyl group in the acylium ion generated, combined with a lower steric bulk of this group could lead to the better performance.

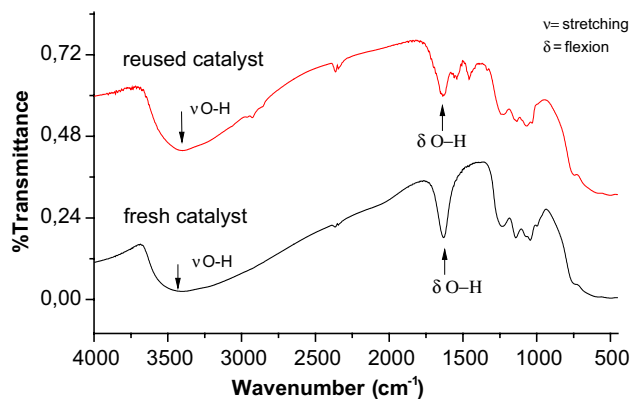


Fig. 3 FT-IR of fresh sulfated zirconia calcined at 550 °C and reused sample

Table 1 Optimization of synthesis of acetyl-3-methylindole **1** using acetic anhydride as acylating reagent

Catalyst	w/w (%)	Ratio substrate/ (CH ₃ CO) ₂ O	Temperature (°C)	Time (h)	Yield (%) (2a)	Yield (%) (3a)	Unreacted reagent (%) (1)
AcOH	100	1:4	82	1	–	–	95
				3	–	–	93
				6	–	–	91
			140	1	–	–	91
				3	–	–	89
				6	–	32	28
ZS	25	1:4	82	1	40	28	27
				1	49	25	10
				0.25	51	46	–
	100	1:2	82	1	46	35	–
				1	53	26	–
				0.25	51	25	10
		140	1:4	1	49	27	9
				0.25	60	30	–
				1	61	30	–

Table 2 Acylation of 3-methylindole **1** and acylating reagents at 140 °C using sulfated zirconia (100% w/w)

Entry	(RCO) ₂ O	Yield (%) (2)	Mp (°C) (2) Found/Lit.*	Yield (%) (3)	Mp (°C) (3) Found/Lit.*
1	(MeCO) ₂ O	60 (2a)	147–148/147–149 [34]	30 (3a)	70–72/66–67 [38]
2	(EtCO) ₂ O	40 (2b)	166–167/161 [35]	39 (3b)	53–54/45 [39]
3	(<i>n</i> -PrCO) ₂ O	60 (2c)	144–146/150–152 [36]	38 (3c)	56–58
4	(PhCO) ₂ O	80 (2d)	140–140.5/138–139 [37]	–	–

3.3 Reusability of Catalyst

Reusability of sulfated zirconia was also studied in the reaction to obtain **3a**. The catalyst recovered after the reaction was employed in new processes. Experiments show that the sulfated zirconia can be reused without significant loss of activity until three reaction cycles (Table 3; Fig. 4).

Table 3 Catalyst reusability studied

Reaction cycle	Yield (%) (2a)	Yield (%) (3a)
1	60	30
2	58	28
3	57	26
4	40	18

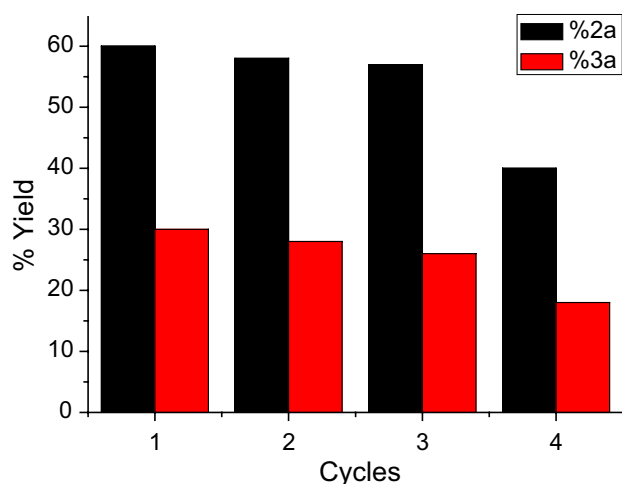


Fig. 4 Catalyst reusability

3.3.1 Physical and Spectral Data for *N*-Butyryl-3-methylindole (**3c**)

Mp: 56–58 °C; $^1\text{H NMR}$ (CDCl_3 , 500 MHz) δ (ppm): 8.48 (d, $J=8.2$ Hz, 1 H, H-7), 7.53 (ddd, $J=7.7$, 1.4, 0.7 Hz, 1 H, H-4), 7.38 (ddd, $J=8.4$, 7.2, 1.4 Hz, 1 H, H-6), 7.32 (td, $J=7.5$, 1.1 Hz, 1 H, H-5), 7.25 (s, 1 H, H-2), 2.88 (t, $J=7.4$ Hz, 2 H, CH_2CO), 2.32 (d, $J=1.4$ Hz, 3 H, CH_3), 1.89 (h, $J=7.4$ Hz, 2 H, $\text{CH}_3\text{CH}_2\text{CH}_2$), 1.10 (t, $J=7.4$ Hz, 3 H, CH_3CH_2). $^{13}\text{C NMR}$ (CDCl_3 , 126 MHz) δ (ppm): 171.11(C=O), 135.92 (C-7a), 131.32(C-3a), 125.10 (C-6), 123.26(C-5), 121.65 (C-2), 118.74(C-4), 118.15 (C-3), 116.65 (C-7), 37.77(CH_2CO), 18.19(CH_2CH_2), 13.83(CH_3CH_2), 9.72(CH_3). Anal. Calcd for $\text{C}_{13}\text{H}_{15}\text{NO}$: C 77.58, H 7.51, N 6.96. Found: C 77.50, H 7.55, N 6.93.

4 Conclusion

In conclusion, an efficient method for the Friedel–Crafts acylation of 3-methylindole with acid anhydrides employing sulfated zirconia as catalyst has been developed. This catalyst shows an excellent catalytic activity, with good conversion and selectivity towards formation of the 2-acylation products, (3-methyl-1*H*-indol-2-yl) ketones. Moreover, the catalyst is air stable and easy to handle with easy work up. In addition, sulfated zirconia can be reused up to three cycles. The characterization of the catalyst performed before and after its use reveals that it did not undergo structural changes. Compared with the use of acetic acid, a better behavior and a smaller production of chemical waste were observed when sulfated zirconia was used, which gives it an additional advantage.

Acknowledgements The authors are grateful to UNLP and CICBA for financial support and Lic. Omar Guaymas for the NMR measurements. D A V and L J M are a holder of CONICET fellowship.

References

1. Sravanthi TV, Manju SL (2016) *Eu J Pharm Sci* 91:1–10
2. Knölker H-J (ed) (2016) *The alkaloids: chemistry and biology*, vol 76. Academic Press, New York 1–340
3. Sundberg RJ (ed) (1996) *Indoles*. Academic, London
4. Stalick WM, Wynne JH, Lloyd CT, Jensen SD, Boson S (2004) *Synthesis* 14:2277–2282
5. Metwally MA, Shaaban S, Abdel-Wahab BF, El-Hiti GA (2009) *Curr Org Chem* 13 (14): 1475–1496
6. Bai G, Li T, Yang Y, Zhang H, Lan X, Li F, Han J, Ma Z, Chen Q, Chen G (2012) *Catal Commun* 29:114–117
7. Guchhait SK, Kashyap M, Kamble H (2011) *J Org Chem* 76:4753–4758
8. Chatterjee A, Biswas KM (1973) *J Org Chem* 38:4002–4004
9. Bergman J, Venemalm L (1987) *Tetrahedron Lett* 28:3741–3744
10. Saxton JE (1952) *J Chem Soc* 3592–3598
11. Ketcha DM, Gribble GW (1985) *J Org Chem* 50:5451–5457
12. Ottoni O, de V F Neder A, Dias AKB, Cruz RPA, Aquino LB (2001). *Org Lett* 3(7):1005–1007
13. Pal M, Batchu VR, Parasuraman K, Yeleswarapu KR (2003) *J Org Chem* 68:6806–6809
14. Pal M, Dakarapu R, Padakanti S (2004) *J Org Chem* 69:2913–2916
15. Cánepa AS, Bravo RD (2004) *J Heterocycl Chem* 41:979–982
16. Cánepa AS, Bravo RD (2006) *J Heterocycl Chem* 43:235–238.
17. Cánepa AS, Sasiambarrena LD, Bravo RD (2008) *Synth Commun* 38:2655–2661
18. Cánepa AS, Bravo RD (2002) *Synth Commun* 32:3675–3680
19. Sasiambarrena LD, Cánepa AS, Bravo RD (2011) *Synth Commun* 41:200–205
20. Sasiambarrena LD, Cánepa AS, Luna CN, Bravo RD (2015) *Tetrahedron Lett* 56(16):2054–2058
21. Sasiambarrena LD, Cánepa AS, Bravo RD (2015) *Tetrahedron Lett* 56(52):7184–7189
22. Sasiambarrena LD, Méndez LJ, Ocsachoque MA, Cánepa AS, Bravo RD, González MG (2010) *Catal Lett* 188:180–186
23. Méndez LJ, Cánepa AS, González MG, Bravo RD (2012) *Tetrahedron Lett* 53(8):688–690
24. Méndez LJ, Cánepa AS, Rimada R, Bravo RD (2013) *Lett Org Chem* 10 (4):240–244
25. Yadav GD, Pujari AA (1999) *Green Chem* 1:69–74
26. Saravanan K, Tyagi B, Shukla RS, Bajaj HC (2016) *Fuel* 165 298–305
27. Patel A, Brahmkhatri V, Singh N (2013) *Renew Energy* 51:227–233
28. Osatiashtiani A, Durndell LJ, Manayil JC, Lee AF, Wilson K (2016) *Green Chem* 18:5529–5535
29. Dar BD, Ahmad N, Patial J, Sharma P, Bindu K, Maity S, Singh B (2014) *J Saudi Chem Soc* 18:464–468
30. Yadav GD, Ajgaonkar NP, Varma A (2012) *J Catal* 292:99–110
31. Reddy BM, Sreekanth PM, Lakshmanan P (2005) *J Mol Catal A Chem* 237:93–100
32. Kuzminska M, Kovalchuk TV, Backov R, Gaigneaux EM (2014) *J Catal* 320:1–8
33. Altass HM, Khder AERS (2016) *J Mol Catal A Chem* 411:138–145.
34. Fuerstner A, Jumbam DN (1993) *J Chem Soc Chem Commun* 2:211–212
35. Jones CD, Suárez T (1972) *J Org Chem* 37(23):3622–3623

36. Grinev AN, Shvedov VI, Krichevskii ES, Romanova OB, Altukhova LB, Kurilo GN, Andreeva NI, Golovina SM, Mashkovskii MD (1984) *Pharm Chem J* 18 (2):94–98
37. Jones CD, Suarez T (1972) *J Org Chem* 37:3622–3623
38. Geissman A (1952) *J Am Chem Soc* 74:3916–3918
39. Odde G (1913) *Gazzetta Chimica Italiana* 43 (II): 198–213