

Nickel selenide thin flms: opto‑electric and thermoelectric properties

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Abstract

Nickel selenide thin flms were synthesized on the amorphous templates utilizing a dip method. In the synthesis process, ascorbic acid acts as complexing agent. Hydrazine hydrate too plays a role of complexing agent and/or catalyst. Structural, optical, microscopic, electrical and thermoelectrical parameters were estimated. The X-ray difraction spectrum of nickel selenide shows polycrystalline nature having hexagonal structure. The absorptivity of nickel selenide flms is highly absorptive with a direct type of transition. The optical gap of NiSe was found to be 2.18 eV. The Urbach energy was found to be 0.275 eV. The ratio of Ni: Se in the thin flm was 1:1. Micrograph of nickel selenide thin flm showing irregular shaped fne grains distributed to the whole surface. The electrical conductance of nickel selenide suggests that the activation energy was found to be 0.024 and 0.331 eV for lower and higher temperature areas, respectively. At 300 K, the carrier concentration is estimated to be 2.58×10^{19} and 3.55×10^{19} at 525 K.

Keywords Thin flms · X-ray difraction · Optical absorption · Thermoelectric power

1 Introduction

Nowadays, transition metal chalcogenides have applied as the absorbent for solar devices to substitute the silicon. They are projected as the most acceptable semiconductor materials for photoelectrochemical cell. It owns outstanding electronic, semiconducting and optical characteristics in the thin film $[1, 2]$ $[1, 2]$ $[1, 2]$ $[1, 2]$. The thickness of such chalcogenides adopts the diferent characteristics. The features of thin flm are dissimilar from bulk sample. It is crucial to choice the proper systems for the preparation of transition metal chalcogenide. [[3–](#page-4-2)[5\]](#page-4-3)

Nickel chalcogenides becomes the novel thin flm substances for the generation of solar power because of its substantial semiconductor and optical properties [\[6\]](#page-4-4). Nickel selenide (NiSe) has been extensively applied in diverse devices such as coating, IR detectors, solar cells, and photoconductors [[7](#page-4-5), [8\]](#page-4-6). Nickel selenide has some stable phases like NiSe, NiSe₂, and Ni₃Se₂ [\[9](#page-4-7)[–11](#page-4-8)]. A different complexing agent like triethanolamine [[12](#page-4-9)], sodium salt of ethylene diammine tetra acetic acid [\[13](#page-4-10)], polyvinyl alcohol [[14](#page-4-11)] was applied to prepare nickel sulphide thin flm. Sharma et al. [[15](#page-4-12)] synthesized nickel selenide thin film using a screen coating technique. The electrical resistivity was found to be $10³$ (Ω.cm) and activation energy around 0.18 eV.

In this research work, nickel selenide thin flms were synthesized by the dip technique at room temperature. Ascorbic acid is utilized as a complexing agent. Structural, optical, microscopic, electrical, and thermoelectrical parameters were estimated.

2 Experimental details

2.1 Templates and chemicals

The deposition was carried on amorphous glass plates of dimensions $26 \times 76 \times 2$ mm. The templates were cleaned by using chromic acid, detergent, and alcohol. Analytical quality chemicals were utilized for the preparation of thin flm. It consists of nickel sulphate octahydrate, ascorbic acid, hydrazine hydrate, sodium sulphite, selenium powder and ammonia. Sodium selenosulphate $(-0.25 M)$ solution is ready by refuxing 2.5 g of selenium powder and 7.5 g sodium sulphite in 100 ml water for 6 h at 363 K. The hot

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solution was cooled and filtered [[16](#page-4-13)]. Nickel selenide thin flm was prepared utilizing dip technique.

2.2 Deposition of NiSe thin flm

To prepare the bath, 10 mL (0.25 M) nickel sulphate octahydrate was poured in 100 mL beaker; other chemicals were used in the following sequence: 2.5 mL (1 M) ascorbic acid, 8 ml (15%) hydrazine hydrate, 10 mL (0.25 M) sodium selenosulphate. The pH of the reactive mixture is 10.95. 10 mL double distilled water is added in the bath. Reactive mixture kept in the ice bath. The non-conducting glass templates maintained vertically in a reaction mixture. The temperature of the reaction mixture surges up to 298 K. The deposition process gets completed after 6 h. The deposited templates were washed with water.

3 Results and discussion

3.1 Growth mechanism

The reaction mechanism of synthesis of NiSe thin flm is shown below;

 Ni^{2+} + Ascorbic acid \rightarrow [Ni – Ascorbate],

$$
\text{Na}_2\text{SeSO}_3 + \text{OH}^- \rightarrow \text{Na}_2\text{SO}_4 + \text{HSe}^-,
$$

 $HSe^- + OH^- \rightarrow H_2O + Se^{-2}$,

[Ni – Ascorbate] + $\text{Se}^{2-} \rightarrow \text{NiSe} + [\text{Ascobate}]^{-2}$.

The nickel ions combine with ascorbic acid. Ascorbic acid binds the nickel ion to prevent the direct reaction between nickel and selenide ions. It plays the role as binding agent. pH of reaction mixture is greater than 7; sodium selenosulphate experiences decay generating free selenide ions. The quantity of nickel ions in the reaction mixture is conserved by ascorbic acid. The temperature of reaction bath upswings progressively; the disintegration of nickel ascorbate complex takes place. Free selenide and nickel ions were combined to form NiSe thin flm. The deposition take place when the ionic product of is greater than solubility product. $[K_{\rm{sp}}=2\times10^{-26}]$ [[17\]](#page-4-14). Hydrazine hydrate is compulsory for the formation of NiSe thin flm. It might be playing a complexing and/or catalytic role in the flm process. It also improves compactness and adherences of the flms [\[18](#page-4-15), [19](#page-4-16)]. The as-grown NiSe flms show grey colour [\[20](#page-4-17)]. The terminal thickness was found to be 0.14 μm for NiSe thin flm.

3.2 X‑ray difraction

The X-ray difraction pattern of the flms was recorded by using Philips-PW-1710, X-ray difractometer, in 2θ range from 10 to 80° using CuK α_1 line (λ = 1.54056 Å). Nickel selenide exhibits wurtzite (hexagonal type), tetragonal and rhombohedral structure. This depends upon synthesis conditions such as the deposition temperature and precursor concentration. The spectra for pure NiSe [JCPDF Card No.75–0610] were used for identifcation purpose. The X-ray difraction spectrum of NiSe thin flm deposited on glass substrate is shown in Fig. [1](#page-1-0).

The XRD pattern shows a large number of peaks indicate that the flm are polycrystalline in nature having hexagonal structure. For NiSe, (100) (102) (110) (103) (112) planes of hexagonal are observed. Using Bragg's equation, the interplanar distance of each peak became estimated. The highest intense reflection at $d = 3.130$ Å can be normally indexed as hexagonal (100). The interplanar distance became 1.984, 1.801, 1.540 and 1.499 Å for (102) (110) (103) and (112) planes, respectively. Lattice parameter of the NiSe thin flm is estimated and found to be *a*=3.614 and *c*=5.129 Å. Umeyama et al. [\[21\]](#page-4-18) reported lattice parameter $a = 3.659$ and $c = 5.352$ Å for NiSe thin flm. Volume of unit cell of NiSe was found to be 66.989 $(A)^3$. The average crystallite size was determined by using Scherer's formula. It was found to be 84.83 nm. The dislocation density of NiSe thin flm was found to be 1.389×10^{14} m⁻². The number of crystallite/unit area was estimated to be 2.293×10^{14} m⁻².

Fig. 1 X-ray difraction pattern of nickel selenide thin flm

Fig. 2 Absorption spectrum of nickel selenide thin flm

Fig. 3 Determination of optical energy of nickel selenide thin flm

3.3 Optical characterization

The absorbance of the NiSe thin flm was measured with respect to wavelength. A Hitachi-330 (Japan) UV–VIS–NIR double-beam spectrophotometer was used for this purpose. The wavelength utilized to obtain the absorption spectra changes from 400 to 800 nm. The absorption spectra were used to calculate band gap, type of optical transition and absorption coefficient. The optical absorption spectra of representative flms are shown in Fig. [2](#page-2-0). Optical density of thin flm decreases with the increase in wavelength. The absorption edge of exponential shape was observed in NiSe thin flm. The absorptive nature of the NiSe thin flm was measured by estimating absorption coefficient. At 400 nm, the absorptive coefficient was estimated to be 9.57×10^4 and 2.85×10^3 cm⁻¹ at 800 nm. It decreases with the increase in wavelength. Higher magnitude of absorption coefficient indicates that NiSe thin flm is highly absorptive. A plot of $(\alpha)^2$ against photon energy should be a straight line whose intercept to the x-axis gives the optical band gap. The graph of $(\alpha)^2$ vs. photon energy for nickel selenide is shown in Fig. [3](#page-2-1). The band gap of NiSe thin flm was found to be

Fig. 4 Determination of band tail energy of nickel selenide thin flm

Fig. 5 EDAX pattern of nickel selenide thin flm

2.18 eV [\[22](#page-4-19)]. The plot of $(\alpha)^2$ versus photon is linear in high energy region indicating direct band to band-type transition. The area close to optical edge is known as Urbach tail. Such tail was observed due to lower crystallinity, imperfections, some non-crystalline part of the materials and dislocations [[23,](#page-4-20) [24\]](#page-4-21). Urbach energy can be found using the slope of the straight line of plotting ln (α) versus photon energy. The reciprocal of the slope gives Urbach energy. For NiSe thin flms, the Urbach energy was found to be 0.275 eV. The graph of $\ln(\alpha)$ versus photon energy is shown in Fig. [4](#page-2-2).

3.4 Compositional and microscopic characterization

The composition of the NiSe thin flm was investigated by using energy-dispersive X-ray analysis technique. The obtained pattern is represented in Fig. [5.](#page-2-3) The quantity of Ni present in the thin flm was 51.36 and selenium 48.64%, respectively. The fraction of nickel to selenium is approximately 1. The microscopic analysis of NiSe thin flm was carried out applying scanning electron microscopy. Figure [6](#page-3-0) indicates the micrograph of NiSe thin flm. The image indicates that the grains are fused with each other. No crack was observed in thin flm. The irregular shaped fne grains were distributed to whole surface. The grains are compactly packed. The average grain dimension is used to estimate

Fig. 6 Scanning electron micrograph of nickel selenide thin flm

by way of applying Contrell's technique. The mean grain dimension as calculated from scanning electron micrograph was applied to be observed in accord with these obtained from Scherer's components in X-ray difraction.

3.5 Electrical characterization

The dark DC conductivity of all samples was measured by using a two-probe method, in temperature range of 300–525 K for heating and cooling cycles. The change in electrical conductance is observed for the frst cycle of heating and cooling. This is because for the frst cycle NiSe thin flm experiences irretrievable medications of non-equilibrium defects. The specifc conductance at room temperature of NiSe was found to be of the order of $10^{-4}(\Omega)$ cm)−1, respectively. The plot of conductivity vs. inverse temperature is shown in Fig. [7](#page-3-1) for the flm. The plot indicates the existence of two kinds of conduction mechanism. The lower-temperature region (300–430 K) was known as intrinsic having smaller slope. Above this temperature, it is known as extrinsic with characterized by a higher slope. The activation energy is calculated using the Arrhenius equation. The activation energies were obtained from

the slope of straight lines of plot. For lower-temperature region, the activation energy was found to be 0.024 and 0.331 eV for higher-temperature area.

3.6 Thermoelectrical power measurement

The temperature range 300 to 525 K is selected to measure electromotive force produced by the thin flm. Thermoelectric power of NiSe thin flm with temperature is signifed in Fig. [8](#page-3-2). The graph indicates the uniform nature of thin flm because of upsurge in thermoelectric power with temperature. The nature of conductivity was also decided by thermoelectric power measurement. In this study, positive terminal is associated with the cold end, so the thin flm confrms p-type conductance. At two extreme points of temperature, thermoelectric power was observed to be 6.0 and 49.5 µV/K. Carrier density is determined using the mathematical form [[25](#page-4-22)]

 $log n = 3/2 log T - 0.005 P + 15.7198.$

Carrier density is estimated for working temperature. At 300 K, the carrier density is obtained to be 2.58×10^{19} cm⁻³ and 3.55×10^{19} cm⁻³ at 525 K. The carrier density of the thin flm became a surge with temperature. Ionized nickel vacancies (Ni^{+2}) are responsible for p-type conductivity in NiSe thin flms. This increase in ionized Ni vacancies with surge temperature enhances the conductivity of the flms. The electrical conductivity rise as shown in Fig. [7](#page-3-1) with respect to temperature can be correlated to carrier density. Conductivity increases as the carrier density increases with temperature.

Fig. 7 Electrical conductivity of nickel selenide thin flm **Fig. 8** Thermoelectrical power of nickel selenide thin flm

4 Conclusions

The preparation of NiSe thin flm on glass templates is carried out by applying the dip method. The hump in the X-ray difraction spectra of NiSe thin flm indicates some amorphous phase that exists in it. The optical characteristics showed that the NiSe thin flm has band tail energy of 0.275 eV. From scanning electron micrograph images, it is suggested that the NiSe thin flm that has irregular shaped grains was attached to each other. The flm shows semiconducting nature having conductance of the order of 10^{-4} (Ω) cm)−1. Tungsten diselenide and molybdenum disulphide have conductivity in the same range as of NiSe. These films have layer structure and better optical properties. The thin flm confrms p-type conductivity.

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